



GCE A LEVEL MARKING SCHEME

SUMMER 2024

**A LEVEL
CHEMISTRY – UNIT 5
1410U50-1**

About this marking scheme

The purpose of this marking scheme is to provide teachers, learners, and other interested parties, with an understanding of the assessment criteria used to assess this specific assessment.

This marking scheme reflects the criteria by which this assessment was marked in a live series and was finalised following detailed discussion at an examiners' conference. A team of qualified examiners were trained specifically in the application of this marking scheme. The aim of the conference was to ensure that the marking scheme was interpreted and applied in the same way by all examiners. It may not be possible, or appropriate, to capture every variation that a candidate may present in their responses within this marking scheme. However, during the training conference, examiners were guided in using their professional judgement to credit alternative valid responses as instructed by the document, and through reviewing exemplar responses.

Without the benefit of participation in the examiners' conference, teachers, learners and other users, may have different views on certain matters of detail or interpretation. Therefore, it is strongly recommended that this marking scheme is used alongside other guidance, such as published exemplar materials or Guidance for Teaching. This marking scheme is final and will not be changed, unless in the event that a clear error is identified, as it reflects the criteria used to assess candidate responses during the live series.

GCE A LEVEL CHEMISTRY UNIT 5**PRACTICAL EXAMINATION****SUMMER 2024 MARK SCHEME****GENERAL INSTRUCTIONS**Recording of marks

Examiners must mark in red ink.

The mark total should be entered onto the grid on the front cover.

Marking rules

All work should be seen to have been marked.

Crossed out responses not replaced should be marked.

Marking abbreviations

The following may be used in marking schemes or in the marking of scripts to indicate reasons for the marks awarded.

cao = correct answer only

ecf = error carried forward

bod = benefit of doubt

EXPERIMENTAL TASK

Skill	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
Teacher-awarded marks	working safely (1) efficient use of time (1)	2			2		2
Results Part A – recording	appropriate table drawn (1) titles and units included (1) appropriate significant figures (1) temperature to 3 sig figs (1 dp) accept volume to 2 or 3 sig figs	2	1		3		3
Results Part A – accuracy and trend	Test 1 final temperature increases from first to second reading (1) final temperature decreases from third to fourth to fifth reading (1) two highest values at 20 cm ³ and 30 cm ³ (1) Test 2 final temperature increases from first to second to third reading (1) final temperature decreases from fourth to fifth reading (1) two highest values at 30 cm ³ and 40 cm ³ (1)		3		3		3

Skill	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
Results Part B – observations	Test 1 X no observable change with either acid (1) Y no observable change with HCl and (white) precipitate forms with H ₂ SO ₄ (1) Z (white) precipitate with both acids (1) Test 2 X (white) precipitate with both acids (1) Y no observable change with either acid (1) Z no observable change with HCl and (white) precipitate forms with H ₂ SO ₄ (1)		3		3		3
Analysis Part A (i) Graph	appropriate label for y-axis (1) appropriate scale for y-axis (1) all five points plotted accurately (2) award (1) for any three/four points plotted accurately		4		4		4
Analysis Part A (ii) Intersection	two appropriate straight lines (1) point of intersection read correctly (1) ecf possible from poorly drawn lines			2	2	2	2

Skill	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
Analysis Part A (iii) Concentration	moles of NaOH = 0.050 mol (1) moles of NaOH = moles of HCl (1) $[\text{HCl}] = \frac{\text{moles of HCl}}{\text{volume of HCl at point of intersection}} \quad (1)$		2	1	3	2	2
Analysis Part A (iv) Enthalpy change	total volume of solution = 50.0 + volume of HCl at point of intersection (1) ΔT from graph (1) n(NaOH) = 0.050 mol (1) correct value for ΔH (1) negative sign for ΔH (1)		3	2	5	3	3

Skill	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
Analysis Part B (v) Identification	Test 1 X contains sodium ions / is sodium nitrate ⇒ no precipitates form because sodium chloride and sodium sulfate / all sodium salts are soluble (1) Y contains calcium ions / is calcium nitrate ⇒ one (white) precipitate because calcium sulfate is insoluble and calcium chloride is soluble (1) Z contains lead ions / is lead(II) nitrate ⇒ two (white) precipitates because lead chloride and lead(II) sulfate are both insoluble (1) if no credit given award (1) for correctly identifying the ions in all three solutions						
	Test 2 X contains lead ions / is lead(II) nitrate ⇒ two (white) precipitates because lead chloride and lead(II) sulfate are both insoluble (1) Y contains sodium ions / is sodium nitrate ⇒ no precipitates form because sodium chloride and sodium sulfate / all sodium salts are soluble (1) Z contains calcium ions / is calcium nitrate ⇒ one (white) precipitate because calcium sulfate is insoluble and calcium chloride is soluble (1) If no credit given award (1) for correctly identifying the ions in all three solutions			3	3		3

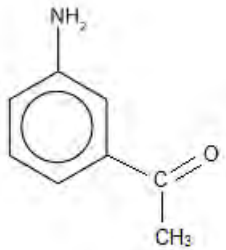
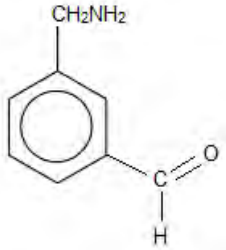
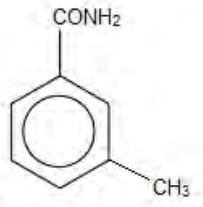
Skill	Marking details	Marks available					
		AO1	AO2	AO3	Total	Maths	Prac
Analysis Part B (vi) Largest enthalpy	award (2) for either of following sodium hydroxide because sodium sulfate is the most soluble of the salts formed sodium hydroxide because calcium sulfate and lead sulfate are insoluble award (1) for sodium hydroxide and reason not linked to solubility e.g. sodium hydroxide because it is the strongest base / Group 1 hydroxides are strong alkalis			2	2		
Total		4	16	10	30	7	25

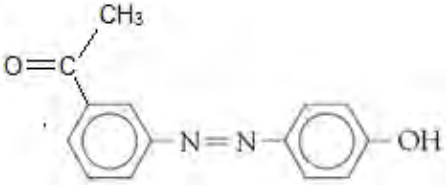
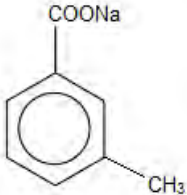
PRACTICAL METHODS AND ANALYSIS TASK

Question				Marking details	Marks available					
					AO1	AO2	AO3	Total	Maths	Prac
1	(a)			filter paper / U-tube of agar soaked in aqueous / saturated KNO_3 accept any suitable electrolyte e.g. KCl	1			1		1
	(b)	(i)		Cu electrode labelled positive			1	1		
		(ii)	I	$\text{Fe} + \text{Cu}^{2+} \rightleftharpoons \text{Fe}^{2+} + \text{Cu}$			1	1		
			II	<u>Fe^{2+} / Fe half-cell</u> award (1) for either of following colour of (pale) green solution becomes deeper (due to increasing concentration of Fe^{2+} ions) size of Fe electrode decreases <u>Cu^{2+} / Cu half-cell</u> award (1) for either of following blue colour of solution becomes paler (due to conversion of Cu^{2+} to Cu / decreasing concentration of Cu^{2+} ions) size of Cu electrode increases / Cu deposited on Cu electrode			2	2		2
				Question 1 total	1	0	4	5	0	3

Question				Marking details	Marks available																									
					AO1	AO2	AO3	Total	Maths	Prac																				
2	(a)			small drops of liquid may fall from the funnel (into the burette) during the titration (leading to a false burette reading)	1			1		1																				
	(b)			award (1) for either of following water does not react with any of the reagents adding water does not change the number of moles of reagents added	1			1		1																				
	(c)			pink	1			1		1																				
	(d)			mass of $\text{FeC}_2\text{O}_4 \cdot x\text{H}_2\text{O} = 3.481 \text{ g}$ (1) <table><tr><td></td><td>Titration 1</td><td>Titration 2</td><td>Titration 3</td><td>Titration 4</td></tr><tr><td>Initial burette reading / cm^3</td><td>1.60</td><td>0.30</td><td>0.80</td><td>1.20</td></tr><tr><td>Final burette reading / cm^3</td><td>29.45</td><td>27.45</td><td>28.00</td><td>28.30</td></tr><tr><td>Titre volume / cm^3</td><td>27.85</td><td>27.15</td><td>27.20</td><td>27.10</td></tr></table> all readings must be given to 2dp award (2) for four correct values award (1) for any two correct values mean titre = 27.15 cm^3 (1)		Titration 1	Titration 2	Titration 3	Titration 4	Initial burette reading / cm^3	1.60	0.30	0.80	1.20	Final burette reading / cm^3	29.45	27.45	28.00	28.30	Titre volume / cm^3	27.85	27.15	27.20	27.10		4		4	2	4
	Titration 1	Titration 2	Titration 3	Titration 4																										
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Question				Marking details	Marks available					
					AO1	AO2	AO3	Total	Maths	Prac
	(e)			$3\text{MnO}_4^- + 5\text{FeC}_2\text{O}_4 + 24\text{H}^+ \rightarrow 3\text{Mn}^{2+} + 5\text{Fe}^{3+} + 10\text{CO}_2 + 12\text{H}_2\text{O}$ award (1) for three correct species added award (1) for correct balancing only when all formulae are correct			2	2		
	(f)			$n(\text{MnO}_4^-) = 0.0215 \times \frac{27.15}{1000} = 0.000584 \text{ mol} \quad (1)$ $n(\text{FeC}_2\text{O}_4 \cdot x\text{H}_2\text{O}) \text{ in } 25 \text{ cm}^3 = 0.000973 \text{ mol} \quad (1)$ $n(\text{FeC}_2\text{O}_4 \cdot x\text{H}_2\text{O}) \text{ in } 500 \text{ cm}^3 = 0.01946 \text{ mol}$ $M_r(\text{FeC}_2\text{O}_4 \cdot x\text{H}_2\text{O}) = \frac{3.481}{0.01946} = 178.9 \quad (1)$ $x = \frac{178.9 - 143.8}{18.02} = 2 \quad (1)$		4		4	4	
				Question 2 total	3	8	2	13	6	7

Question			Marking details	Marks available					
				AO1	AO2	AO3	Total	Maths	Prac
3	(a)		compound A  tests 1 and 2 show that a carbonyl group is present and that it is a ketone / not an aldehyde tests 4 and 5 shows that it has an aromatic amine group						
			compound B  tests 1 and 2 show that a carbonyl group is present and that it is an aldehyde test 4 shows that it has an aliphatic amine group			6	6		6
			compound C  test 1 shows that it has no carbonyl group test 3 shows that it has an amide group (which is hydrolysed giving ammonia gas)						
			for each compound award (2) for correct structure and clear reasoning for one of the groups award (1) for one correct group with clear reasoning award (1) for correct structure with no clear reasoning						

Question			Marking details	Marks available					
				AO1	AO2	AO3	Total	Maths	Prac
	(b)	(i)			1		1		
		(ii)			1		1		
	(c)	(i)	award (1) for correct conversion of units $28^{\circ}\text{C} \Rightarrow 301\text{ K}$ $1\text{ atm} \Rightarrow 1.01 \times 10^5\text{ Pa}$ $75.2\text{ cm}^3 \Rightarrow 75.2 \times 10^{-6}\text{ m}^3$ $\text{moles of gas} = \frac{1.01 \times 10^5 \times 75.2 \times 10^{-6}}{8.31 \times 301} = 0.00304\text{ mol} \quad (1)$	1	1		2	2	
		(ii)	moles of compound B = 0.00304 mol mass of compound B = $0.00304 \times 135.09 = 0.411 \quad (1)$ concentration = $\frac{0.411}{0.2} = 2.06\text{ g dm}^{-3} \quad (1)$ ecf possible from part (i)			2	2	2	2
			Question 3 total	1	3	8	12	4	8

A2 UNIT 5: PRACTICAL EXAMINATION

SUMMARY OF ASSESSMENT OBJECTIVES

	Question	AO1	AO2	AO3	TOTAL MARK	MATHS	PRAC
Experimental Task	Total	4	16	10	30	7	25
Practical Methods and Analysis Task	1	1	0	4	5	0	3
	2	3	8	2	13	6	7
	3	1	3	8	12	4	8
TOTAL		9	27	24	60	17	43